



ALLIANCE™

(<https://www.globalseafood.org>).



 Responsibility

Chlorine an effective disinfectant in aquaculture

1 November 2008

By Claude E. Boyd, Ph.D.

Chlorine compounds can be toxic to culture species and workers



Chlorine is an effective treatment for disinfecting hatchery tanks.

Chlorination is practiced widely in aquaculture for disinfection of holding and rearing tanks for fish. Chlorine compounds are applied to puddles of water in the bottoms of empty ponds to eliminate wild fish before refilling ponds. Chlorination also has been used to kill disease organisms or their vectors in water supplied to hatcheries and in shrimp ponds before stocking.

Chlorine compounds can be toxic to culture species, but they also can be toxic to workers. Those adopting chlorination as a management practice should have knowledge of the behavior of chlorine compounds in order to use them effectively and safely.

Chlorine sources

The three common sources of chlorine are chlorine gas, sodium hypochlorite or household bleach, and calcium hypochlorite, commonly called high-test hypochlorite or HTH.

Chlorine gas often is used for disinfection of municipal water supplies. It is delivered in railroad tank cars or tank trucks, or generated on site. However, the chlorine is highly toxic to humans and was used in World War I against troops. Chlorine gas is therefore not safe for use at aquaculture facilities.

A comprehensive solution for the wild seafood supply chain.

- Crew rights
- Food safety
- Environmental responsibility

Best Seafood Practices

LEARN MORE

(<https://bspcertification.org/>).

Sodium hypochlorite is a dilute, liquid solution, but because of its low percentage of active ingredient, it is not widely used in aquaculture. A granular material containing about 65 percent active ingredient, calcium hypochlorite is preferred for aquacultural purposes.

Reactions in water

The active ingredients in chlorination products are chlorine gas, hypochlorous acid (HOCl) and hydrochlorite ions (OCl-) that result from the dissociation of HOCl. These forms are called free chlorine residuals. In terms of disinfecting power, chlorine gas and HOCl are about 100 times more potent than OCl-.

When a chlorination product is applied to water, it dissolves and chlorine speciation occurs based on pH. At pH below 2, chlorine gas is dominant. Between pH 2 and 6, HOCl is the only type. As pH increases beyond 6, OCl- appears. At a pH of about 7.5, HOCl and OCl- occur in equal proportions, but at higher pH, OCl- is the most abundant chlorine.

Application rates for chlorination products should be increased as pH rises. For high pH values, the dose rates for achieving 1 mg/l of free chlorine residual using HTH with 65 percent active ingredient are as follows: pH 7.0, 1.54 mg/l; pH 7.5, 2.34 mg/l; pH 8.0, 4.71 mg/l; pH 8.5, 11.80 mg/l; pH 9.0, 29.20 mg/l.

Free chlorine residuals are strong oxidants that oxidize organic matter, nitrite and other reduced substances. Free chlorine residuals are reduced to nontoxic chloride in these reactions, and their disinfecting power is lost.

Chlorine species also react with ammonia to form chloramines that have less disinfecting power than free chlorine residuals. Chloramines are called combined chlorine residuals. Some vendors sell chloramines for use as disinfectants for aquaculture, but these products are not as effective as calcium hypochlorite.

Free chlorine residuals quickly decompose to chloride and water in the presence of light. In outdoor systems, the residual action of chlorination lasts only a few hours, or at most a few days.

Applications

Chlorination usually is applied to provide a free chlorine residual of 1-3 mg/l. The amount of calcium hypochlorite necessary to provide this concentration depends upon pH (Table 1), as discussed above. It also is influenced by the abundance of plankton and bacteria, and the concentrations of dissolved and suspended organic matter, ammonia nitrogen, nitrite and other reduced substances.

Boyd, Effect of pH on hypochlorous acid:hydrochlorite ion ratios, Table 1

pH	Hypochlorous Acid: Hydrochlorite	Free Chlorine Residual Needed for Disinfecting Power of 1 mg/l at pH 7 Ion Ratio
6.0	32.00	0.78
6.5	10.00	0.84
7.0	3.20	1.00
7.5	1.00	1.52

8.0	0.30	3.06
8.5	0.10	7.68
9.0	0.03	19.00

Table 1. Effect of pH on hypochlorous acid:hydrochlorite ion ratios.

The best way to determine the required dosage is to measure the chlorine demand. This is the amount of chlorine product necessary to react with living organisms and reduced substances in water. The treatment dose should be equal to the chlorine demand plus enough of the product to provide 1-3 mg/l free chlorine residual.

Since the chlorine demand test is difficult to conduct, most farmers use a dose proven effective previously. In Southeast Asia, shrimp farmers often treat ponds before stocking postlarvae with 20-30 mg/l of calcium hypochlorite. However, products provided by different vendors can vary in strength, and a wide range in concentration of free chlorine residuals probably is applied to ponds.

Chlorine effects

When done correctly, chlorination can kill possible wild vectors of disease in ponds. However, a study conducted at Auburn University revealed that small doses of chlorine added to channel catfish ponds were rendered ineffective immediately by extraneous reactions with organic matter. Chlorination also kills plankton and benthos, leaving little natural food available when postlarvae are stocked. Shrimp ponds should be fertilized a few days after chlorination to encourage the regrowth of natural food organisms before introducing postlarvae.

Chlorine solutions sprayed over the bottoms of empty ponds during pond preparation are not likely to be highly effective in eliminating disease organisms harbored in the soil from the previous crop. Soil contains a large amount of organic matter relative to water, and treatment of moist soil with up to 500 mg/kg active chlorine from high-test hypochlorite did not reduce microbial abundance in a research trial. Treatment of the pond water usually is more effective.

Some farmers think that chlorination of water or bottom soil reduces organic matter concentrations. Although chlorine does oxidize organic matter, at the concentrations used in aquaculture, it only reacts with highly reactive functional groups on organic molecules and does not cause a significant reduction in organic matter concentrations in water or bottom soil.

Chlorinated water from municipal water supplies often is used to fill tanks or aquariums for holding aquatic organisms. Although this water initially may have sufficient residual chlorine to severely stress or kill aquatic animals, it can be held a few days until the chlorine residuals have dissipated through exposure to light and extraneous reactions.

Alternatively, the chlorine can be removed immediately by application of sodium thiosulfate. It requires about 7 mg/l sodium thiosulfate to remove 1 mg/l free residual chlorine. Sodium thiosulfate is not appreciably toxic to fish and other organisms and is often applied according to previous experience.

Safety

Chlorination should be applied cautiously. Chlorine gas is dangerous to use at aquaculture operations because of the possibility of accidental release and subsequent injury or death to workers. Sodium or calcium hypochlorite can generate toxic chlorine gas if put into highly acidic water. Spills of chlorine products into ponds or natural waters can destroy aquatic life. Workers should wear protective clothing and masks with suitable air filters when working with concentrated sources of chlorine.

(Editor's Note: This article was originally published in the November/December 2008 print edition of the Global Aquaculture Advocate.)

Author



CLAUDE E. BOYD, PH.D.

Department of Fisheries and Allied Aquacultures
Auburn University
Alabama 36849 USA

boydce1@auburn.edu (<mailto:boydce1@auburn.edu>)

Copyright © 2024 Global Seafood Alliance

All rights reserved.